



SDI FINAL EVALUATION FORM 1.1

PART 1:

Journal Name:	British Journal of Applied Science & Technology
Manuscript Number:	MS: 2012/BJAST/1901
Title of the Manuscript:	PHOTOCATALYTIC DEGRADATION OF CHLOROBENZENE IN AQUEOUS SOLUTIONS – COMPARISON OF TIO ₂ WITH PERIWINKLE AND SNAIL SHELLS' POWDER AS PHOTOCATALYST

PART 2:

FINAL EVALUATOR'S comments on revised paper (if any)	Authors' response to final evaluator's comments
<p>This manuscript was promoted evidently thanks for the authors' hard work; it still, however, needs revision before it can be accepted for possible publication. Detail comments are:</p> <ol style="list-style-type: none">Figures 3.1 and 3.2 are too dirty for use as scientific paper figures, so please used them as supporting information.Table 2, I believe that same SSA and PSA were used when H₂O₂ was added, so the different specific areas of SSA+H₂O₂ and PSA+H₂O₂ could be attributed to the reaction between SSA, PSA and H₂O₂, please clarify it.Fig. 3.3-3.5, please clarify irradiation conditions, under dark ? or sunlight irradiation? or other conditions?The results shown in Fig. 3.6 are difficult to understand. The increase of CB degradation between the first 1-4 hours of reaction is reasonable; however, further increasing the reaction time resulted in the decrease of degradation efficiency, indicating an increase of CB concentration and lose of activity, why?Other mistakes: Mortal and pestle → Mortar and pestle?	<ol style="list-style-type: none">Figures 3.1 and 3.2 which are scans of the FT-IR spectra had been moved to appendix as requestedPure SSA has lower specific area (16 m²/g) when compared to pure PSA (21 m²/g) as a result of the larger particle size of SSA. On addition of H₂O₂ the surface area of both powders increase to 29 m²/g and 38 m²/g for SSA and PSA respectively. We believe that this astronomical increase may not be unrelated with the reactive nature of H₂O₂. The H₂O₂ may have acted on the powders reducing their particle size leading to increase in their surface area. This is however, an assumption since there was no way to test the reaction mechanism.Since this set of experiments were principally aimed at determining the optimum reaction conditions, they were conducted under sunlight irradiationThe Figure is self explanatory. After the first 1-4 hours of reaction time, the degradation rate for pure PSA and SSA dropped minimally and remained relatively constant as time increases while the degradation rate dropped sharply for TiO₂ and H₂O₂ modified SSA and PSA before remaining constant as time increases. For all the catalysts, the degradation rate decreases as time increases beyond 4 hours.This is a typographical error. It had been corrected